

Writing Ionic Equation of the Reactions of Acid

- Write the chemical equation of the reaction.
- List down all the ion presence below the equation.
- Cancel the ions that do not change in the reaction.
- Rewrite the equation with only the ions that have changed in the reaction.

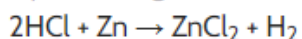
Example

- Write the ionic equations for the following reactions:
 - Acids + Zinc
 - Acids + Metal Carbonate
 - Acids + Base Oxide
 - Acids + Alkali

Answer

- Acids + Zinc
 - We can use any acid for the reaction, as it will not affect the ionic equation at the end.
 - In this case, let's use hydrochloric acid for the chemical equation.

Step 1: Writing the Chemical Equation



Step 2: List down the ions



Step 3: Cancel the ions that do not change

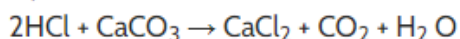


Step 4: Rewrite the equation

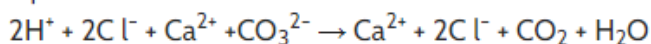


- Acids + Metal Carbonate
 - Let's use hydrochloric acid and calcium carbonate

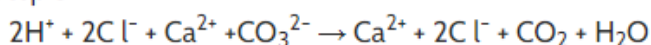
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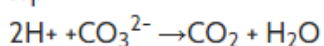
Step 2:



Step 3:



Step 4:

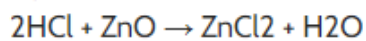




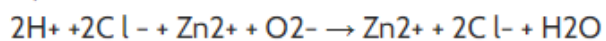
- **Acids + Base Oxide**

- **Let's use Hydrochloric acid and Zinc Oxide**

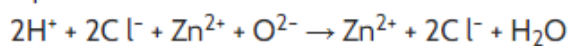
Step 1:



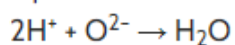
Step 2:



Step 3:

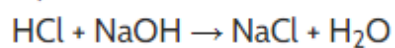


Step 4:

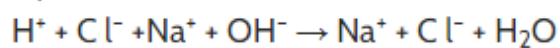


- **Acids + Alkali**

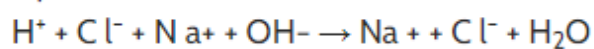
Step 1:



Step 2:



Step 3:



Step 4:

